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Low-Temperature Isolation of An Azidophosphenium Cation

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In Table 1 of this Communication, the last line was misprinted. The table with the correct Q_{CT}^{tot} values is depicted here.

Table 1: Calculated partial charges [e] and charge transfer Q_{CT}^{tot} [e] in an isolated ion pair of **1**, **2a**–**c**, and **4**[14b] along with partial charges of the $[(Me_3Si)_2N{=}P{-}X]^+$ ion.[a]

	1	**2a**	**2b**	**2c**	**4**
$q(P_{salt})$	1.05	1.19	1.21	1.19	1.41
$q(N_{amino,salt})$	−1.65	−1.54	−1.51	−1.52	−1.51
$q(X_{salt})$[a]	−0.31	−0.21	−0.20	−0.17	−0.33
$q(P_{cat})$	1.24	1.26	1.20	1.20	1.37
$q(N_{amino,cat})$	−1.49	−1.52	−1.46	−1.46	−1.47
$q(X_{cat})$[a]	−0.18	−0.18	−0.18	−0.17	−0.31[e]
Q_{ct}^{tot}[b]	0.66[c]	0.19[d]	0.12[d]	0.12[d]	0.07[d]

[a] Compound **1** was formally considered as the salt $[(Me_3Si)_2N{=}P{-}Cl]^+[Cl]^-$. **1** and **2a**–**c** X = Cl, **4** X = N_3. [b] Q_{ct}^{tot} = charge transfer with respect to the $[(Me_3Si)_2N{=}P{-}X]^{n+}$ ion (X = Cl for **1**, **2a**–**c** and X = N_3 for **4**), thus $Q_{cation} = 1 - Q_{CT}^{tot}$. [c] $Q_{CT}^{tot} = q(Cl^-)$. [d] $Q_{CT}^{tot} = 1 + \Sigma q(A_i)$ with the A_i atom of the anion. [e] $q(N_{azide,salt}) = -0.72$ versus $q(N_{azide,cat}) = -0.70$.

The editorial office apologizes for this mistake.

